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C-10 SF Na
STRONG ACID CATION EXCHANGE RESIN
(Designed for use in water softening applications)

Product Description

C 11 ODQ(Na) resin is a high-capacity, conventional gel polystyrene strong acid cation exchange resin designed for use in residential or industrial water softening equipment. Cation resin in sodium form removes hardness ions such as calcium and magnesium by replacing them with sodium. When the resin bed is exhausted the hardness ions begin to pass through the bed. Functionality is returned by regeneration with concentrated sodium or potassium chloride solution. The capacity obtained depends largely on the amount of salt used in the regeneration. Typically 12-15 lbs. of chemical per ft3 is used to obtain maximum capacity of up to 36,000 grains per ft3 (claim not tested/certified by WQA).

C11ODQ-Na resin is rigorously treated before shipping to meet all NSF-44 standards requirements.

Typical Physical, Chemical & Operating Characteristics

Polymer Structure	Polystyrene 10% cross-linked with Divinylbenzene
Physical Form and Appearance	Amber spherical beads
Whole Bead Count	90% Min.
Functional Group	Polystyrene Sulfonate
Ionic Form (as shipped)	Na+
Shipping Weight, approx.	850g/l (53 lbs./ft3)
Mesh Size (U.S. Std.)	16-50
Moisture Retention, Na+ form	40-47%
Swelling, Na+->H+	5%
Total Capacity in sodium form	2.1 meq/ml (claim not tested/certified by WQA)
pH Range, Stability	0-14

CHEMICAL AND THERMAL STABILITY

CI 1 ODQ(Na) resin is insoluble in dilute or moderately concentrated acids, alkalies, and in all common solvents. However, exposure to >0.5 ppm of free chlorine, "hypochlorite" ions, or other strong oxidizing agents over long periods of time will eventually break down the crosslinking. Temperature over 30 °C (85 °F) will accelerate the oxidation. This will tend to increase the moisture retention of the resin, decreasing it's mechanical strength, as well as generating small amounts of extractable breakdown products. Like all conventional Polystyrene sulfonated resins, it is thermally stable to higher than 160 °C (320 °F) in the alkali (for instance, sodium) or alkaline earth (calcium and magnesium) salt forms. The free acid form tends to hydrolyze in water temperatures appreciably higher than 130 °C (270 °F) thereby losing capacity, as the functional groups are gradually replaced by hydroxyl groups.